



# Cardinal Heenan Catholic High School





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## GCSE Chemistry – Half Term 3 CPL Schedule

This schedule is designed to help you to focus your revision each week and weekend for GCSE chemistry. It is based on the most common areas for improvement for the class based on the December mocks.

**You will have a weekly quiz to check this revision.**

Date	Topic	Revision YouTube clips	Practice Exam Questions
Friday 17 <sup>th</sup> January 2025	<b>Ionic Bonding and Electrolysis Solutions</b>  Old revision guide pages 16-17 and 43-43 New revision guide page 29 and 58-59 <b>Problematic Question</b> Describe what happens to calcium atoms and chlorine atoms when the ionic compound calcium chloride is formed	<a href="https://www.youtube.com/watch?v=XbDtmORzKO8">https://www.youtube.com/watch?v=XbDtmORzKO8</a>  <a href="https://www.youtube.com/watch?v=6WjCVi4roA">https://www.youtube.com/watch?v=6WjCVi4roA</a>	
Friday 24 <sup>th</sup> January 2025	<b>Chemical cells</b>  Old revision guide pages 58-59 New revision guide page 64 <b>Problematic Question</b> The voltage produced by a chemical cell depends on the concentration of the electrolyte solution. Plan an experiment to investigate how the voltage produced by a chemical cell varies with the concentration of the electrolyte solution. The following substances are available: - The metal electrodes in table 3 - 1.0mol/dm <sup>3</sup> sodium chloride solution - Pure water	<a href="https://www.youtube.com/watch?v=riikUBLFBJs">https://www.youtube.com/watch?v=riikUBLFBJs</a>	
Friday 31 <sup>st</sup> January 2025	<b>Percentage by Mass and Volume of Gas Equation</b>  Old revision guide page 33 New revision guide page 46  <b>Problematic Question</b> Fe <sub>2</sub> O <sub>3</sub> reacts with carbon to produce carbon dioxide. The equation for the reaction is $2\text{Fe}_2\text{O}_{3(s)} + 3\text{C}_{(s)} \rightarrow 4\text{Fe}_{(s)} + 3\text{CO}_{2(g)}$ Calculate the volume of carbon dioxide gas at room temperature and pressure that is produced from 40.0Kg of Fe <sub>2</sub> O <sub>3</sub> using excess carbon	<a href="https://www.youtube.com/watch?v=VaXQVoI3gtc">https://www.youtube.com/watch?v=VaXQVoI3gtc</a>  <a href="https://www.youtube.com/watch?v=tYE-1nywlfFs&amp;t=32s">https://www.youtube.com/watch?v=tYE-1nywlfFs&amp;t=32s</a>  <a href="https://www.youtube.com/watch?v=OzutTI0sYd0">https://www.youtube.com/watch?v=OzutTI0sYd0</a>	
Friday 7 <sup>th</sup> February 2025	<b>Properties of Simple Covalent Molecules and Bond Energies</b>  Old revision guide pages 20-21 and 58 New revision guide page 32 and 63  <b>Problematic Question</b> Explain why propane has a low boiling point	<a href="https://www.youtube.com/watch?v=u_KROUaZfKY">https://www.youtube.com/watch?v=u_KROUaZfKY</a>  <a href="https://www.youtube.com/watch?v=eExCBkp4jB4">https://www.youtube.com/watch?v=eExCBkp4jB4</a>  <a href="https://www.youtube.com/watch?v=PdValXAVUOc">https://www.youtube.com/watch?v=PdValXAVUOc</a>	




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Friday 14 <sup>th</sup> February 2025	<b>Calculating the Rates of a Chemical Reaction</b>  Old revision guide pages 61-61 New revision guide page 71  <b>Problematic Question</b> Determine the mean rate of reaction for small calcium carbonate lumps between 20s to 105s. Give the unit. You need to calculate the gradient of the line here, using rise over run	<a href="https://www.youtube.com/watch?v=UkrBJ6-uGFA&amp;list=PL9IouNCPbCxW8AN0t0py7LaKdKSwfL3fP">https://www.youtube.com/watch?v=UkrBJ6-uGFA&amp;list=PL9IouNCPbCxW8AN0t0py7LaKdKSwfL3fP</a>	
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